

Introduction to General Chemistry CHM 1025C

Chapter Test-2, Ch 4,5,11

NOTE: Failure to Follow Directions
Failure put the Units in
Failure to properly round

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**Failure to show all equations
Failure to show all math work,
= POINTS OFF**

PRINT YOUR NAME ON THE LINE: _____

(1 pt ea)

Your start time on this test _____

Your finish time on this test: _____

Time it took you to do this test: _____

I. Elemental Knowledge FILL IN THE BLANKS OR CIRCLE THE CORRECT ANS.

- 1 An Element is the same as a Compound? True False
2. Name and give the chemical symbols for the three most abundant elements in the earth crust, ocean and atmosphere
A-Name. _____ A-Sym _____
B-Name. _____ B-Sym _____
C-Name. _____ C-Sym. _____
3. Name and give the chemical symbols for the three most abundant elements in the human body:
A-Name. _____ A-Sym _____
B-Name. _____ B-Sym _____
C-Name. _____ C-Sym. _____
4. A given compound always has the same chemical composition means:
A. It will always be a solid, liquid or gas
B. It will always have the same relative number of atoms
C. It will always have the same exact number of atoms
D. It will obey Boyles Law
5. A compound is composed of A. An Element C. Protons, neutrons and electrons
B. Atoms of two or more elements D. None of the above
6. Electrons have a _____ charge
7. Write the names of the Protons, Electrons and Neutrons in order of mass
Largest Mass _____
Smaller Mass _____
Smallest Mass _____

8. Rutherford's experiment demonstrated that the atomic structure was not that of plum pudding. He used a _____

_____ particle and it had a _____ charge

9 What determines chemical behavior?

A. Electrons

C. Neutrons

B. Protons

D. None of the above

10. An atom with different number of neutrons is called

A. Isotope

D. Metalloid

B. Atomic Number

E. Allotrope

C. Mass Number

F. None of the above

11. The number of protons in the nucleus of an atom is also called it's _____

12. The number of protons and neutrons in an atom is also called it's _____

13. A Nitrogen atom may contain 7 protons, 7 electrons and 7 neutrons. Write out the symbol for this Nitrogen atom showing this data:

14. The periodic table has atoms in order of _____

15. Give another name for at least 2 of the columns in the periodic table that we have dealt with:

A. _____

B. _____

16. Vertical columns are also called: _____

17. Name [Name and Symbol] at least two diatomic molecules

A-Name. _____ A-Sym _____

B-Name. _____ B-Sym _____

18. When we take an electron away from a neutral atom we form a

19. Salt water conducts electricity, sugar in water does not. Why?

II Multiple Choice – Compounds

1. When we say we're Naming a BINARY COMPOUND, we mean: _____
A. A Compound that contains one molecule
B. A Substance that contains one element
C. A Compound that contains two elements
D. A Substance that contains one molecule
E. None of the above
2. Binary Compounds can contain: _____
A. A Metal
B. A Metal and a Non-Metal
C. A Non-Metal
D. None of the above
5. To become an Anion, an element must: _____
A. Loose an electron
B. Gain an electron
C. Loose a neutron
D. None of the above
6. When naming a binary ionic compound: _____
A. The Anion is named first
B. The Cation is named first
C. The most complex part of the compound is named first
D. None of the above
7. How many electrons can be in one orbital? _____
A. 1
B. 2
C. 3
D. 4
E. None of the above.
8. A Chemical Formula shows _____
A. The number of protons, neutrons and electrons in the compound
B. The types and number of atoms in the substance
C. The types and number of atoms in the element
D. The types and number of atoms in the compound
E. None of the above
9. What determines chemical behavior? _____
A. Electrons
B. Protons
C. Neutrons
D. None of the above

III. Name the following Compounds or Give the formula

1. Cobalt (II) Bromide _____
2. AlBr_3 _____
3. NO _____
4. PCl_5 _____
5. K_2S _____
6. H_2SO_4 _____
7. HI _____
8. Sodium Sulfate _____
9. Potassium Phosphate _____
10. Magnesium Perchlorate _____
11. Tin (IV) Chromate _____
12. Lead (II) Carbonate _____
13. Copper (I) Chloride _____
14. Iron (III) Sulfate _____
15. Ammonium Nitrate _____

Summary:

Section	# of Quest	Pts Ea	Tot Pts	# Right	Total Pts
1. Elemental Knowledge	19	2	38	_____	_____
2. Mult Choice	9	2	18	_____	_____
3. Name Compounds	15	3	45	_____	_____
Total			101		

How do you rate this test from 1 to 10 _____

1 = Very East, can do it with my eyes closed

10= Very Very Difficult, could not do any of the problems